be seen, for example, in the strikingly different redox potentials of secondary and primary hydroxylamines. Primary hydroxylamines have redox potentials in the 300 mV range (see Fig.3 of Tamilmani et al., 2003, DuPont Electronic Technology, http://www.ekctech.com/images/feauture-stories/MRS-Interaction%20between%20ceria%20and%20Hydroxylamine.pdf, enclosed), near that of the intracellular reducing potential (e.g. Sies, et al., 1977, Euro J Biochem 72, 301-7, abstract attached), whereas the cyclic secondary hydroxylamines of Krisha et al. provide redox potentials ranging from 722 to 960 mV (see, e.g. Krishna et al., 1992, PNAS USA 89, 5537-41, attached).

Krishna (1998) describes use of dozens of different compounds, but every one is a similar cyclic secondary nitroxide (and the corresponding hydroxylamines and amines). Krisha (1998) provides no suggestion or motivation to deviate from his teachings and employ a structurally and functionally distinct class of hydroxylamines, particularly since Krisha (1998) repeatedly reports that variation in redox potential across his reagents showed no significant correlation between protection and redox potentials (e.g. Krishna, 1998, at p.3488, col.2, line 41-45).

The Examiner is invited to call the undersigned if he would like to amend the claims to clarify the foregoing or seeks further clarification of the claim language.

We petition for and authorize charging our Deposit Account No.19-0750 all necessary extensions of time. The Commissioner is authorized to charge any fees or credit any overcharges relating to this communication to our Dep. Acct. No.19-0750 (order B00-001-2).

Respectfully submitted,
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Encl. Krishna et al. (1998, Journal of Medicinal Chemistry 41(18):3477-92)
Sies et al. (1977, Euro J Biochem 72, 301-7, abstract)
Krishna et al. (1992, PNAS USA 89, 5537-41).
Tamilmani et al. (2003, Dupont Electronic Technology, 6 p.)

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